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rwilliams1. Terms in this set (7) actual yield.
the amount of product that forms when a
reaction is carried out in the laboratory.
excess reagent.

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Stoichiometry The study of quantitative
relationships between the amounts of
reactants used and amounts of products
formed by a chemical reaction is called
stoichiometry. Stoichiometry is based on the
law of conservation of mass.

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Hunde: Zeichnen Sie süße und lustige Hunde und Welpen. Buch zum Zeichnen und Färbefür Kinder und Anfänger, blau-weiße Abdeckung mit Hunden etc.

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Stoichiometry 1. $1 \text{ ZnI}_2 \rightarrow 1 \text{ Zn} + 1 \text{ I}_2$ How many grams of iodine will you produce if you begin your rxn with 56.7g of zinc iodide? 56.7 g 1 mole 1 mole 253.8 g 319.19 g 1 mole 1 mole = 45.08 grams of Iodine 2. $1 \text{ Pb}_2(\text{CO}_3)_3 + 6 \text{ K} \rightarrow 2 \text{ Pb} + 3 \text{ K}_2\text{CO}_3$

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 $1 \text{ Zn} + 1 \text{ I}_2$ How many grams of iodine will you produce if you begin your rxn with 56.7g of zinc iodide? $56.7 \text{ g} \cdot \frac{1 \text{ mole}}{253.8 \text{ g}} \cdot \frac{1 \text{ mole}}{1 \text{ mole}} = 45.08$ grams of Iodine 2. $1 \text{ Pb}_2(\text{CO}_3)_3 + 6 \text{ K}_2\text{CO}_3$
 $\text{Pb} + 3 \text{ K}_2\text{CO}_3$

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56.7g of

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